**Concentration of salt-battery**

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**Evaluation:**

The data of our experiment was really different from what we had expected. By adding more salt each time to tap water, we expected that the voltage and current will increase. After the experiment, we had a conclusion of; the amount doesn’t affect the voltage and current, and no stable change.

One of the error that caused our whole experiment to change was we left our two metals; copper and iron, in the water for two days, and the salt water caused the rust on the metals to wash off and contaminate the tap water. Because of this, we had to change our experiment, so instead of dumping out the water, we diluted the solution-changing the concentration. The weakness of this experiment was we didn’t know that the salt water will wash off the rust on the metals, which changed our experiment and made a sudden change. The strength of this experiment is even with the sudden change, we still kept the other variables the same; temperature of water, amount of water, distance between the metals, size of the metals.

Does the amount of salt change the current or voltage of a "salt water battery"? The best way for this experiment to answer our question is that if we knew there would be a sudden change, we can start off with more salt dissolved in the tap water, and dilute it each time.

The answer that we have concluded from our experiment is that the amount of salt isn’t the big cause or main reason for the voltage and current to have a big difference. I think in order for this experiment to answer our question, we should do it again, and this time, add more salt and then dilute it.

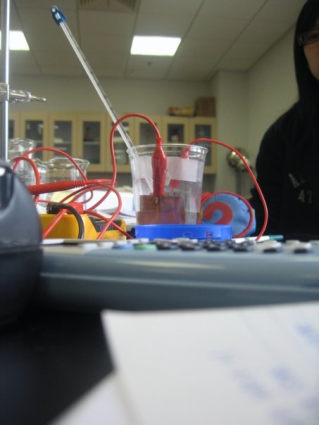


Figure 2: This is what happened after leaving our metals in our solution for two days.(rust)

Figure 1: this is how we set up our experiment.

